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$q_1$  and  $q_2$  = two electric charges  
 $r$  = distance between two electric charges  
 $F$  = Force

2.  $V = \frac{W}{Q}; W = V \times Q; Q = \frac{W}{V}$

$V = p.d.$ ,  $W$  = work done,  $Q$  = Quantity of charge transferred

3.  $V = R \times I; R = \frac{V}{I}; I = \frac{V}{R}$

$V = p.d.$ ;  $R$  = Resistance,  $I$  = current.

4.  $R = \frac{\rho \times l}{A}; \rho = \frac{R \times A}{l}$

$R$  = Resistance;  $l$  = length;  $A$  = Area of cross section;  $\rho$  = rho, a constant known as resistivity

5. Series combination  $R = R_1 + R_2 + R_3 \dots + R_n$

6. Parallel combination  $\frac{1}{R} = \frac{1}{R_1} + \frac{1}{R_2} + \frac{1}{R_3} \dots + \frac{1}{R_n}$

For equal resistances

$R_s = nR$  (For series connection)

$R_p = \frac{R}{n}$  (For parallel connection)

$\frac{R_s}{R_p} = n^2$

$R_s$  = Effective resistance in series

$R_p$  = Effective resistance in parallel

$n$  = number of resistors

$R$  = Resistance of each resistor

7.  $P = \frac{W}{t}$ ; Power =  $\frac{\text{work}}{\text{time}} = \frac{\text{Energy consumed}}{\text{Time}}$

8.  $W = V \times I \times t$ ; Power = potential difference  $\times$  current  $\times$  time

$\Rightarrow (W = I^2 R t) \Rightarrow (W = \frac{V^2 t}{R})$

9.  $P = V \times I$ ; Power = potential difference  $\times$  current

10.  $P = I^2 \times R$ ; Power = (current)<sup>2</sup>  $\times$  resistance

11.  $P = \frac{V^2}{R}$ ; Power =  $\frac{(\text{potential difference})^2}{\text{resistance}}$

12. Electric energy =  $P \times t$ ; electric energy = power  $\times$  time



Multiplying co-efficient of  $x$  with  $\frac{1}{2}$  i.e.,  $\frac{1}{2}(-2) = -1$   
 Now adding  $(-1)^2$  on both sides of the equation (i) we have

$$x^2 - 2x + (-1)^2 = 195 + (-1)^2$$

$$(x - 1)^2 = 195 + 1$$

$$(x - 1)^2 = 196$$

Taking square root of both sides,

$$\sqrt{(x - 1)^2} = \sqrt{196}$$

$$x - 1 = \pm 14$$

Either,

$$x - 1 = 14$$

$$\text{or } x - 1 = -14$$

$$x = 14 + 1$$

$$\text{or } x = -14 + 1$$

$$x = 15$$

$$\text{or } x = -13$$

$$\text{Solution Set} = \{-13, 15\}$$

(vii)  $-x^2 + \frac{15}{2} = \frac{7}{2}x$

**Solution:**

$$-x^2 + \frac{15}{2} = \frac{7}{2}x$$

Multiplying by 2, both sides

$$-2x^2 + 15 = 7x$$

$$15 = 7x + 2x^2$$

$$\Rightarrow 2x^2 + 7x = 15$$

Dividing both sides by 2,

$$x^2 + \frac{7}{2}x = \frac{15}{2} \dots\dots\dots (i)$$

Multiplying co-efficient of  $x$  with  $\frac{1}{2}$  i.e.,  $\frac{1}{2}(\frac{7}{2}) = \frac{7}{4}$

Now adding  $(\frac{7}{4})^2$  on both sides of the equation (i) we have

$$x^2 + \frac{7}{2}x + (\frac{7}{4})^2 = \frac{15}{2} + (\frac{7}{4})^2$$

$$(x + \frac{7}{4})^2 = \frac{15}{2} + \frac{49}{16}$$

$$(x + \frac{7}{4})^2 = \frac{120 + 49}{16}$$

$$(x + \frac{7}{4})^2 = \frac{169}{16}$$

Name: Mrs. Simpson  
Date: April 3, 2018

Skill or Standard for focus / Objective: RI.2.5  
I can determine and evaluate the structure that an author uses to organize a text

### Chronological Order

When a writer writes in chronological order, they put events in the order that they happened.

Also called Sequence of Events

• biographies about people's lives are usually written in chronological order.  
• A writer might use chronological order to tell about historical events.

Ways to show chronological order:

- lists
- timelines
- storyboards
- flowcharts

Using chronological order is important because it helps to show the relationship between events which helps us to understand cause and effect of those events

Understanding the structure of a text can help us to determine a writer's purpose for writing the text

Structure refers to the way a writer creates a text.

Chronological order is a way of organizing a text so that events are listed in the order that they happen.

### Signal Words

- ordinal numbers like first, second, third, etc.
- next • before
- last • later
- after • then

### Example Question:

How does the structure of the text contribute to the text's meaning?

(In other words, what is the writer trying to tell us?)

Summary: Chronological order is a type of structure where a writer gives information in the order that it happened. This helps us to understand how events affect or influence each other.

### REVIEW QUESTION:

#### MULTIPLE CHOICE QUESTIONS

- The process of gaseous exchange involves:  
(a) Breakdown of C-H bonds to yield energy.  
(b) Physical movements that take air in and out of body.  
(c) Getting oxygen from the air and removing carbon dioxide.  
(d) Transport of oxygen by the blood to different parts of the body.
- Most of the gaseous exchange in a leaf occurs through:  
(a) Stomata (b) General surface (c) Cuticle (d) Lenticels
- How many bronchi are there in the air passageway?  
(a) One (b) Two (c) Many (d) None
- Where does the gaseous exchange occur in humans?  
(a) Pharynx (b) Trachea (c) Bronchi (d) Alveoli
- Which structure actively helps in taking the air out of lungs?  
(a) Nasal cavity (b) Bronchus (c) Bronchiole (d) Diaphragm
- The primary chemical stimulus for breathing is the concentration of:  
(a) Carbon dioxide in blood (b) Oxygen in blood  
(c) Carbon dioxide in muscles (d) Oxygen in muscles
- Point out the FALSE statement about respiration:  
(a) Gases can easily pass through the walls of the alveoli  
(b) Gas exchange in lungs is very efficient because lungs provide large surface area  
(c) In emphysema the walls of alveoli break and there is more surface area  
(d) Dust particles can damage the lung by irritating the inner alveoli surface.
- A disease involving the breakdown of air sacs of the lungs is:  
(a) Pneumonia (b) Bronchitis (c) Asthma (d) Emphysema
- Which process does not occur in the nasal cavity?  
(a) Trapping of large dust particles (b) Humidification of the inhaled air  
(c) Warming of the inhaled air (d) Exchange of gases
- What type of blood vessels surrounds the alveoli?  
(a) Artery (b) Arteriole (c) Capillary (d) Vein

### ANSWER KEY

Q.No	Ans	Q.No	Ans	Q.No	Ans	Q.No	Ans
1	c	2	a	3	b	4	d
5	a	6	c	7	d	8	d
9	d	10	c				

- (a) N<sub>2</sub> (b) O<sub>2</sub> (c) Cl<sub>2</sub> (d) S
- Which gas is used to manufacture sulphuric acid?  
(a) N<sub>2</sub> (b) O<sub>2</sub> (c) Cl<sub>2</sub> (d) S
- Equilibrium constant has no unit when number of moles of reactants and products are:  
(a) same (b) different  
(c) both a & b (d) none
- For reactions having large Kc value, the reaction proceeds to:  
(a) completion (b) equilibrium state  
(c) back ward (d) None
- The characteristics of reversible reactions are the following except:  
(a) products never react back to form reactants  
(b) they never complete  
(c) they proceed in both ways  
(d) they have a double arrow between reactants and products
- In the blue lake, the reaction  $\text{CaCO}_3 \rightleftharpoons \text{Ca}^{2+} + \text{CO}_3^{2-}$  goes to completion because:  
(a) of high temperature  
(b) it is more stable than CaCO<sub>3</sub>  
(c) CO<sub>2</sub> escapes continuously  
(d) CaO is not dissociated
- For the reaction,  $2\text{A}_2 + \text{B}_2 \rightleftharpoons 2\text{A}_2\text{B}$ , the expression for the equilibrium constant is:  
(a)  $\frac{[\text{A}_2][\text{B}]}{[\text{A}_2\text{B}]}$  (b)  $\frac{[\text{A}_2]^2[\text{B}]}{[\text{A}_2\text{B}]^2}$
- (a) [C] (b) [C]<sup>2</sup>  
(a) [A] [B] (b) [A]<sup>2</sup> [B]<sup>2</sup>
- When a system is at equilibrium state:  
(a) the concentrations of reactants and products becomes equal  
(b) the opposing reaction (forward and reverse) stop  
(c) the rate of the forward reaction becomes very low  
(d) the rate of the forward and reverse reactions become equal
- Which one of the following statements is not correct about active mass?  
(a) rate of reaction is directly proportional to active mass  
(b) active mass is taken in molar concentration  
(c) active mass is represented by square brackets  
(d) active mass means total mass of substances
- When the magnitude of Kc is very large it indicates:  
(a) reaction mixture consists of almost all products  
(b) reaction mixture has almost all reactants  
(c) reaction has not gone to completion  
(d) reaction mixture has negligible products
- When the magnitude of Kc is very small it indicates:  
(a) equilibrium will never establish

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